













CC13 Groups in the Periodic Table







CC13a Group 1

Step	Learning outcome	Had a look	Nearly there	Nailed it!
 7 th	Explain the classification of alkali metals, halogens and noble gases, into groups in the periodic table.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 6 th	Describe the main physical properties of alkali metals.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 6 th	Describe the reactions of lithium, sodium and potassium with water.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Write word, balanced and H ionic equations (including state symbols) for the reactions of alkali metals.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Describe the pattern of reactivity of the alkali metals.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 10 th	Explain how the electronic configurations of the atoms of alkali metals affect their reactivity.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>





CC13b Group 7

Step	Learning outcome	Had a look	Nearly there	Nailed it!
 3 rd	Recall the appearance of chlorine, bromine and iodine at room temperature.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Describe the trends in colour, melting point and boiling point of chlorine, bromine and iodine down the group, and use these to predict physical properties of other halogens.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 5 th	Describe the chemical test for chlorine gas.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Describe the trends in the reactions of halogens with metals, and use this to predict reactions of other halogens.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Write word and balanced chemical equations, including state symbols, for the reactions of halogens with metals.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 6 th	Describe hydrogen halides and their chemical properties.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC13c Halogen reactivity




Step	Learning outcome	Had a look	Nearly there	Nailed it!
 6 th	Describe the relative reactivity of halogens.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Explain how the reactivity of halogens can be worked out from displacement reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Write balanced chemical equations, including state symbols, for the displacement reactions of halogens.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 10 th	H Explain how displacement reactions are examples of redox reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	H Write ionic equations, including state symbols, for displacement reactions of halogens.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 10 th	Explain the order of reactivity of halogens (using electronic configurations).	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC13d Group 0






Step	Learning outcome	Had a look	Nearly there	Nailed it!
 8 th	Explain why noble gases are chemically inert by referring to their electronic configuration.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 3 rd	Describe uses of noble gases linked with their properties.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 6 th	Describe the trends in the physical properties of the noble gases.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Use trends in physical properties to predict the physical properties of other noble gases.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC14 Rates of Reaction







CC14a Rates of reaction

Step	Learning outcome	Had a look	Nearly there	Nailed it!
 6 th	Describe different changes that can occur as a reaction proceeds.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 7 th	Suggest different experimental methods to investigate rates of reaction (e.g. measurements of mass of reactants against time, volume of gas released against time, concentration of reactant or product against time).	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 7 th	Use graphs of changes (in mass, volume or concentration of reactant or product) against time, to interpret what is happening during reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC14b Factors affecting reaction rates




Step	Learning outcome	Had a look	Nearly there	Nailed it!
 8 th	Explain what has to happen for reactions to take place.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Explain why changes in the energy of particles affect rates of reaction.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Explain why changes in the frequency of collisions between particles affect the rate of reaction.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Explain why changes in temperature, concentration, surface area and pressure affect the rate of reaction (surface area for solids, pressure for gases only).	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Describe ways of speeding up or slowing down chemical reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC14c Catalysts and activation energy






Step	Learning outcome	Had a look	Nearly there	Nailed it!
 6 th	Describe what a catalyst does.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 7 th	Explain how catalysts are useful.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Explain what the activation energy of a reaction is.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Explain how catalysts speed up chemical reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 7 th	Describe what enzymes are.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 6 th	Name one or more examples of enzymes.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC15 Heat energy changes in chemical reactions

CC15a Exothermic and endothermic reactions

Step	Learning outcome	Had a look	Nearly there	Nailed it!
 4 th	Recall some examples of exothermic and endothermic changes.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 7 th	Describe how heat changes in solution may be determined.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 8 th	Describe the differences between exothermic and endothermic changes.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

CC15b Energy changes in reactions

Step	Learning outcome	Had a look	Nearly there	Nailed it!
 9 th	Describe exothermic and endothermic reactions in terms of energy changes when bonds are broken and formed.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 1 st	 Use bond energies to calculate energy changes in reactions.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 9 th	Explain the meaning of activation energy.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
 10 th	Draw and label reaction profiles.	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>