CC3 Atomic Structure

CC3a Structure of an atom

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8th	Describe how Dalton's ideas about atoms have changed.	☐	☐	☐
8th	Describe how the subatomic particles are arranged in an atom.	☐	☐	☐
8th	Explain how atoms of different elements are different.	☐	☐	☐
7th	Recall the charges and relative masses of the three subatomic particles.	☐	☐	☐
8th	Explain why all atoms have no overall charge.	☐	☐	☐
8th	Describe how the size of an atom compares to the size of its nucleus.	☐	☐	☐

CC3b Atomic number and mass number

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7th	State where most of the mass of an atom is found.	☐	☐	☐
7th	State the meaning of atomic number.	☐	☐	☐
7th	State the meaning of mass number.	☐	☐	☐
8th	Describe how the atoms of different elements vary.	☐	☐	☐
8th	State the number of electrons in an atom from its atomic number.	☐	☐	☐
8th	Calculate the numbers of protons, neutrons and electrons using atomic and mass numbers.	☐	☐	☐

CC3c Isotopes

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7th	State what is meant by an isotope.	☐	☐	☐
7th	Identify isotopes from information about the structure of atoms.	☐	☐	☐
8th	Calculate the numbers of protons, neutrons and electrons using atomic numbers and mass numbers.	☐	☐	☐
9th	Explain why the relative atomic mass of many elements is not a whole number.	☐	☐	☐
10th	H Calculate the relative atomic mass of an element from the relative masses and abundances of its isotopes.	☐	☐	☐